Name\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_

ACTIVITY SERIES OF METALS.

1. Why is gold a suitable metal for making jewellery? (2 marks)

***Non reactive***

***attractive***

***malleable and ductile***

***strong Any 2=1mark. explanation =1 mark.***

2. Explain why calcium is not found alone as an element in nature. (2 marks)

***far too reactive in nature, combining with H2O and O2***

3a.Why can zinc be used as a coating to stop iron from rusting? (2 marks)

***Forms a protective layer of zinc oxide which stops the iron from rusting.***

4. Explain why tin cans will rust when scratched but galvanised iron will not.

(2 marks)

***Zinc forms a protective coat of zinc oxide. Tin does not, it is just a physical protection so if scratched the iron is exposed and rusts.***

5. Why does sodium have to be stored under paraffin oil? (1 mark)

***Very reactive when exposed to water/water vapour so it must be covered.***

6. When copper is used as a decorative item it is covered with a laquer after being

polished. What is the reason for this? (1 mark)

***Copper will tarnish so, if you want to avoid constant polishing, cover it with a laquer***.

7. Sacrificial anodes are used on aluminium boats.

a. Explain what is meant by a sacrificial anode. (1 mark)

***A metal that is more reactive than aluminium is attached to the boat.***

b. Describe the purpose of the anode and how it works. (1 mark)

***The anode will react before the aluminium therefore preventing the aluminium from oxidising.***

8. If copper metal is placed in a solution containing silver, what will happen? (1 mark)

***Copper will displace the silver form the solution so it will seem to disappear and the silver will form metal crystals.***

References = 2 marks. Therefore they must be correct.